Chemistry
Solutions Webquest

Name_____
Date _____ Period ____

Click on the following website to go to your first site

http://www.solubilityofthings.com/basics/

1. What are the two parts of a solution? _____
2. Define a solute:
3. Define solvent:
4. Define solubility:
5. What is Equilibrium in chemistry?

Click on the following website to go to your second site...

http://www.chem.purdue.edu/gchelp/solutions/character.html

6. The concentration of a solution represents the _____ of _____ _____ in a unit amount of _____ or of solution.
7. Concentrated solutions have a _____ amount of solute.
8. Diluted solutions have a _____ amount of solute.
9. Describe an unsaturated solution:
10. Describe a saturated solution:
11. What happens when you add more solute to an unsaturated solution?
12. What happens when you add more solute to a saturated solution?

Click the following link to go to your third site…

http://ga.water.usgs.gov/edu/solvent.html

13. Why is water called a "universal solvent"?

14. What makes water so good at dissolving other things?

15. How do water and our kidneys work together?

16. Why does salt dissolve in water?

Click the following link to go to your fourth site…
http://www.solubilityofthings.com/basics/factors_affecting_solubility.php

17. What are the five factors that affect solubility?

1. _____
2. _____
3. _____
4. _____
5. _____

18. How is solubility affected by temperature?

19. Define polarity:

20. How does polarity affect solubility?

21. What is the aphorism used by chemists to describe polarity?

22. How does the pressure of a gas affect solubility?

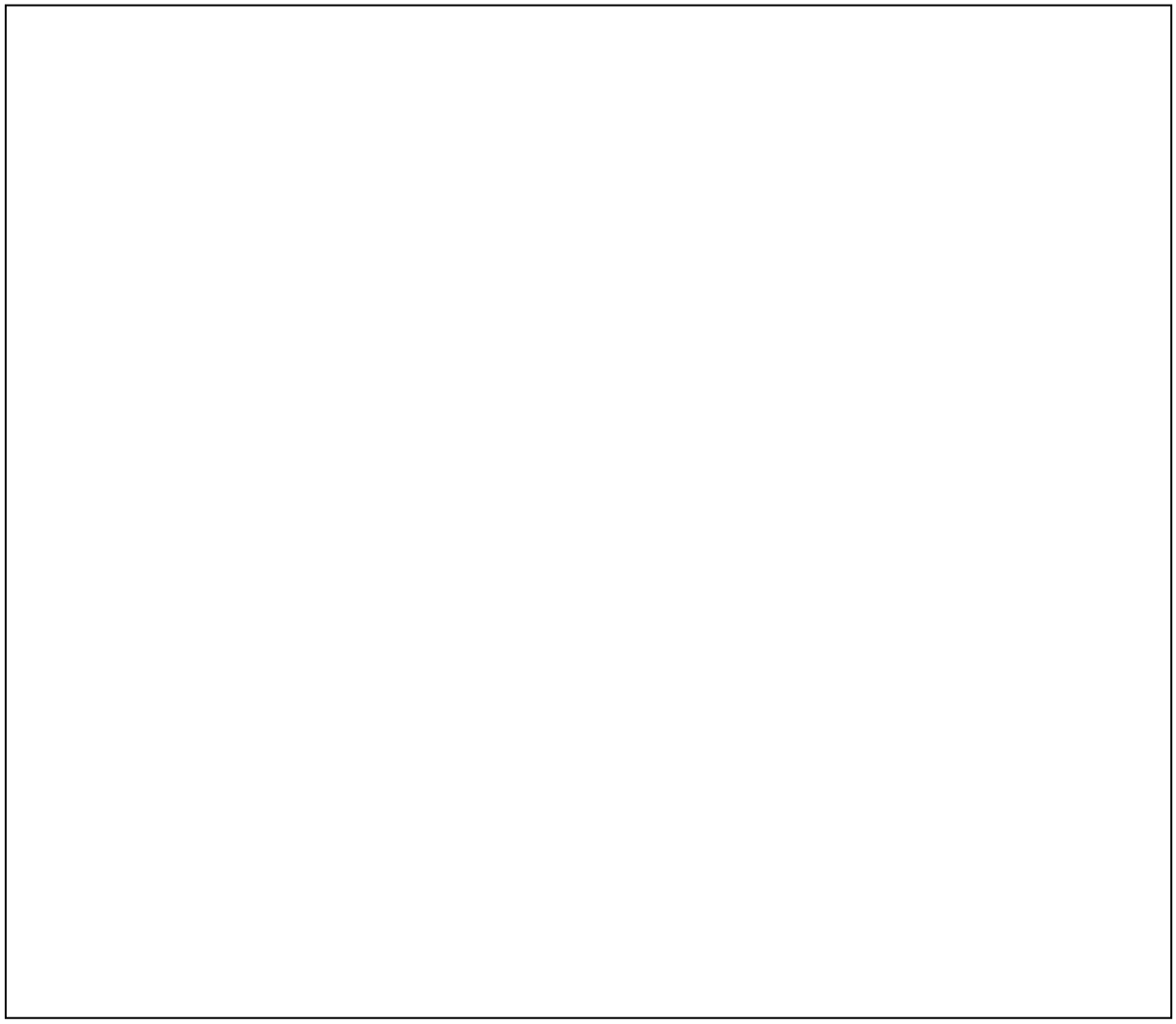
23. How does molecular size affect solubility?

24. How does stirring affect solubility?

PART 2: Click the following link to go to your fifth site…
http://www.kentchemistry.com/links/Kinetics/SolubilityCurves.htm

25. What are solubility curves used for?

26. Draw and label the solubility curve you see on the website below.



27. On the line=_____ (_____ can _____ hold anymore _____)

28. Below the line=_____ (_____ hold more _____)

29. Above the line=_____ (holding _____ _____ than it should, _____ condition)

Click on the link below to go to your sixth site:
https://www.chem.purdue.edu/gchelp/howtosolveit/Solutions/concentrations.html

30. There are a _____ of ways to express the relative amounts of solute and solvent in solution.

31. List four ways to express concentration:

1. 3.

2. 4.

32. What TWO things must you have to compute the following...

Molarity	_____	_____
Molality	_____	_____
Percent by Mass	_____	_____

33. Does temperature affect the following computations?

Molarity _____

Molality _____

34. What is the difference between MOLARITY and MOLALITY?

35. What is the formula for MOLARITY?

36. What is the unit that is used to for MOLARITY?

37. What symbol represents MOLARITY?

38. What is the difference between MOLES and MOLARITY

Click on the following website to go to your seventh site...

http://www.glencoe.com/qe/scienceOLC.php?qi=6277

You must do all 12 problems. When you complete the quiz click on the check button at the Bottom of the page. Record the results of the quiz below.

Number Right _____ Number Wrong _____ Score% _____