

## Balancing Equations V

- 1) \_\_\_\_\_ Al + \_\_\_\_\_ ZnCl<sub>2</sub> → \_\_\_\_\_ AlCl<sub>3</sub> + \_\_\_\_\_ Zn
- 2) \_\_\_\_\_ As<sub>2</sub>O<sub>5</sub> + \_\_\_\_\_ H<sub>2</sub>O → \_\_\_\_\_ H<sub>3</sub>AsO<sub>4</sub>
- 3) \_\_\_\_\_ NH<sub>3</sub> + \_\_\_\_\_ O<sub>2</sub> → \_\_\_\_\_ NO + \_\_\_\_\_ H<sub>2</sub>O
- 4) \_\_\_\_\_ C<sub>3</sub>H<sub>6</sub> + \_\_\_\_\_ O<sub>2</sub> → \_\_\_\_\_ CO<sub>2</sub> + \_\_\_\_\_ H<sub>2</sub>O
- 5) \_\_\_\_\_ NaClO<sub>3</sub> → \_\_\_\_\_ NaCl + \_\_\_\_\_ O<sub>2</sub>
- 6) \_\_\_\_\_ Ca + \_\_\_\_\_ O<sub>2</sub> → \_\_\_\_\_ CaO
- 7) \_\_\_\_\_ LiOH + \_\_\_\_\_ H<sub>3</sub>PO<sub>4</sub> → \_\_\_\_\_ Li<sub>3</sub>PO<sub>4</sub> + \_\_\_\_\_ H<sub>2</sub>O
- 8) \_\_\_\_\_ C<sub>6</sub>H<sub>10</sub> + \_\_\_\_\_ O<sub>2</sub> → \_\_\_\_\_ CO<sub>2</sub> + \_\_\_\_\_ H<sub>2</sub>O
- 9) \_\_\_\_\_ LiOH + \_\_\_\_\_ Fe(NO<sub>3</sub>)<sub>3</sub> → \_\_\_\_\_ LiNO<sub>3</sub> + \_\_\_\_\_ Fe(OH)<sub>3</sub>
- 10) \_\_\_\_\_ P<sub>4</sub> + \_\_\_\_\_ O<sub>2</sub> → \_\_\_\_\_ P<sub>2</sub>O<sub>3</sub>
- 11) \_\_\_\_\_ Fe(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>)<sub>2</sub> + \_\_\_\_\_ BeCl<sub>2</sub> → \_\_\_\_\_ FeCl<sub>2</sub> + \_\_\_\_\_ Be(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>)<sub>2</sub>
- 12) \_\_\_\_\_ Pb(NO<sub>3</sub>)<sub>2</sub> + \_\_\_\_\_ NH<sub>4</sub>Cl → \_\_\_\_\_ PbCl<sub>2</sub> + \_\_\_\_\_ NH<sub>4</sub>NO<sub>3</sub>
- 13) \_\_\_\_\_ MgCO<sub>3</sub> + \_\_\_\_\_ H<sub>3</sub>PO<sub>4</sub> → \_\_\_\_\_ Mg<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub> + \_\_\_\_\_ H<sub>2</sub>O + \_\_\_\_\_ CO<sub>2</sub>
- 14) \_\_\_\_\_ AgF + \_\_\_\_\_ Fe<sub>2</sub>(CO<sub>3</sub>)<sub>3</sub> → \_\_\_\_\_ FeF<sub>3</sub> + \_\_\_\_\_ Ag<sub>2</sub>CO<sub>3</sub>
- 15) \_\_\_\_\_ AgNO<sub>3</sub> + \_\_\_\_\_ AlCl<sub>3</sub> → \_\_\_\_\_ AgCl + \_\_\_\_\_ Al(NO<sub>3</sub>)<sub>3</sub>
- 16) \_\_\_\_\_ Al + \_\_\_\_\_ Fe<sub>3</sub>N<sub>2</sub> → \_\_\_\_\_ AlN + \_\_\_\_\_ Fe
- 17) \_\_\_\_\_ Na + \_\_\_\_\_ Cl<sub>2</sub> → \_\_\_\_\_ NaCl
- 18) \_\_\_\_\_ H<sub>2</sub>O<sub>2</sub> → \_\_\_\_\_ O<sub>2</sub> + \_\_\_\_\_ H<sub>2</sub>O
- 19) \_\_\_\_\_ C<sub>12</sub>H<sub>22</sub>O<sub>11</sub> + \_\_\_\_\_ O<sub>2</sub> → \_\_\_\_\_ CO<sub>2</sub> + \_\_\_\_\_ H<sub>2</sub>O
- 20) \_\_\_\_\_ NaOH + \_\_\_\_\_ H<sub>3</sub>PO<sub>4</sub> → \_\_\_\_\_ Na<sub>3</sub>PO<sub>4</sub> + \_\_\_\_\_ H<sub>2</sub>O
- 21) \_\_\_\_\_ LiClO<sub>3</sub> → \_\_\_\_\_ LiCl + \_\_\_\_\_ O<sub>2</sub>
- 22) \_\_\_\_\_ Al<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub> + \_\_\_\_\_ BaBr<sub>2</sub> → \_\_\_\_\_ AlBr<sub>3</sub> + \_\_\_\_\_ BaSO<sub>4</sub>
- 23) \_\_\_\_\_ (NH<sub>4</sub>)<sub>3</sub>PO<sub>4</sub> + \_\_\_\_\_ Pb(NO<sub>3</sub>)<sub>2</sub> → \_\_\_\_\_ NH<sub>4</sub>NO<sub>3</sub> + \_\_\_\_\_ Pb<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub>
- 24) \_\_\_\_\_ C<sub>7</sub>H<sub>16</sub> + \_\_\_\_\_ O<sub>2</sub> → \_\_\_\_\_ H<sub>2</sub>O + \_\_\_\_\_ CO<sub>2</sub>