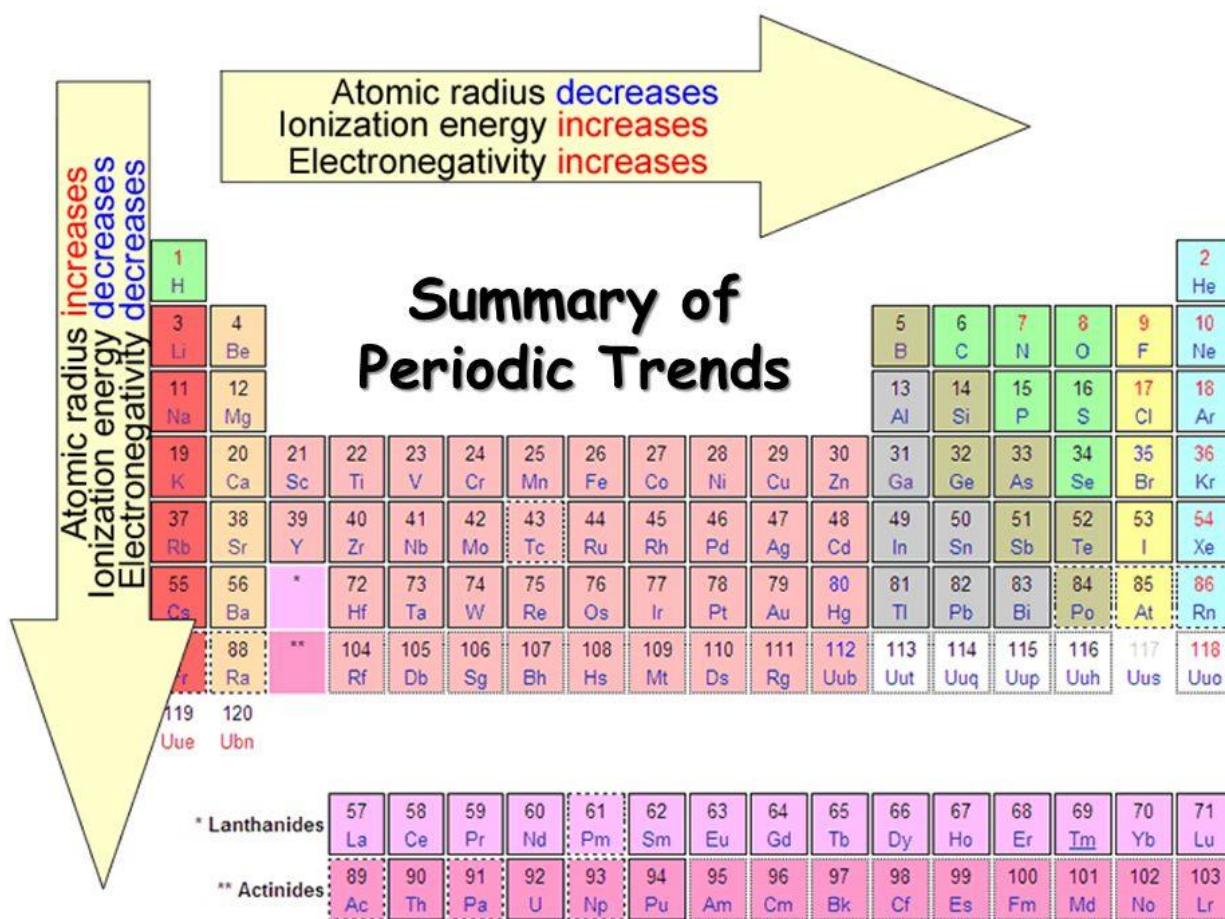


Atomic Radius – indicates size of an atom

Electronegativity is the ability of an atom of an element to attract electrons when the atom is in a compound.

Ionization energy - The energy required to remove an electron from an atom.

Use the following table as a review of atomic size, electronegativity and ionization energy.



For more information about atomic radii, ionization energy, and electronegativity, pull your work from your notes. A thorough explanation of the trends can be found:

<https://www.youtube.com/watch?v=ViIDHZb9iE8>

Multiple Choice

Identify the choice that best completes the statement or answers the question.

1. What is the oxidation number for Oxygen?

- (a) +2 (b) +1 (c) -2 (d) -1

2. Find the oxidation number for Calcium.

- (a) +2 (b) +3 (c) -2 (d) -1

3. What is the oxidation number for Iodine?

- (a) +2 (b) -3 (c) -2 (d) -1

4. How many valence electrons does Neon have?

- (a) 2 (b) 6 (c) 7 (d) 8

5. Find the number of valence electrons for Potassium.

- (a) 1 (b) 2 (c) 3 (d) 4

6. How many valence electrons does Nitrogen have?

- (a) 2 (b) 3 (c) 4 (d) 5

7. Which of the following has the largest atomic radius?

- (a) N (b) C (c) B (d) O

8. Which of the following has the largest atomic size?

- (a) Al (b) B (c) In (d) Ti

9. Which of the following has the smallest atomic size?

- (a) K (b) Na (c) Li (d) Cs

10. Which of the following has the largest ionization energy?

- (a) Be (b) Ca (c) Ra (d) Sr

11. Which of the following has the smallest ionization energy?

- (a) K (b) Ca (c) Mn (d) Co

12. Which of the following has the largest ionization energy?

- (a) F (b) Cl (c) I (d) Br

13. Which of the following has the greatest electronegativity?

- (a) Sr (b) Rb (c) Mo (d) Zr

14. Which of the following has the smallest electronegativity?

- (a) Si (b) Sn (c) Pb (d) Ge

15. Which of the following has the largest electronegativity?

- (a) Li (b) B (c) Be (d) C