Oxidation Numbers of Common Monatomic Ions

1+

Cu^{+}	copper(I)
H^{+}	hydrogen
Pb^{+}	lead(I)
Li^{+}	lithium
Hg^{+}	mercury(I)
K^{+}	potassium
Ag^{+}	silver(I)
Na^{+}	sodium

3+

Al^{3+}	aluminum
Cr^{3+}	chromium(III)
Co^{3+}	cobalt(III)
Fe^{3+}	iron(III)
Pb^{3+}	lead(III)
Mn^{3+}	manganese(III)
Ni^{3+}	nickel(III)

4+

Pb^{4+}	lead(IV)
Mn^{4+}	manganese(IV)
Sn^{4+}	tin(IV)

2+

Ba^{2+}	barium
Cd^{2+}	cadmium
Ca^{2+}	calcium
Cr^{2+}	chromium(II)
Co^{2+}	cobalt(II)
Cu^{2+}	copper(II)
Fe^{2+}	iron(II)
Pb^{2+}	lead(II)
Mg^{2+}	magnesium
Mn^{2+}	manganese(II)
Hg^{2+}	mercury(II)
Ni^{2+}	nickel(II)
Ag^{2+}	silver(II)
Sr^{2+}	strontium
Sn^{2+}	tin(II)
Zn^{2+}	zinc

6+

Cr^{6+}	chromium(VI)
Mn^{6+}	manganese(VI)

1-

Br^{-}	bromide
Cl^{-}	chloride
F^{-}	fluoride
H^{-}	hydride
I^{-}	iodide

2-

O^{2-}	oxide
S^{2-}	sulfide

3-

N^{3-}	nitride
P^{3-}	phosphide

Charges of Common Polyatomic Ions

1-

CH_3COO^{-}	acetate
BrO_3^{-}	bromate
ClO_4^{-}	perchlorate
ClO_3^{-}	chlorate
ClO_2^{-}	chlorite
ClO^{-}	hypochlorite
CN^{-}	cyanide
FO_3^{-}	fluorate
FO_2^{-}	fluorite
OH^{-}	hydroxide
IO_4^{-}	periodate
IO_3^{-}	iodate
NO_3^{-}	nitrate
NO_2^{-}	nitrite
MnO_4^{-}	permanganate

2-

CO_3^{2-}	carbonate
CrO_4^{2-}	chromate
$Cr_2O_7^{2-}$	dichromate
$C_2O_4^{2-}$	oxalate
O_2^{2-}	peroxide
SiO_3^{2-}	silicate
SO_4^{2-}	sulfate
SO_3^{2-}	sulfite
$S_2O_3^{2-}$	thiosulfate

3-

AsO_4^{3-}	arsenate
PO_4^{3-}	phosphate

1+

NH_4^{+}	ammonium

Electronegativity

Fluorine	3.98
Oxygen	3.44
Chlorine	3.16
Nitrogen	3.04
Bromine	2.96
Iodine	2.66
Sulfur	2.58
Selenium	2.55
Carbon	2.55
Hydrogen	2.20
Phosphorus	2.19
Boron	2.04
Silicon	1.90
Aluminum	1.61