

Chemistry 2018-2019

9 weeks	Unit	Title	Length
1 st	1	Scientific Method and Measurement	3 weeks
	2	Matter and Energy	6 weeks
2 nd	3	Atomic Theory and Structure Nuclear Chemistry	4 weeks
	4	Periodic Table	2 weeks
3 rd	5	Bonding and Nomenclature	4 weeks
	6	Chemical equations and Reactions	3 weeks
4 th	7	Stoichiometry	3 weeks
	8	Solutions Acid/Base	4 weeks

Unit 1: Scientific Method and Measurement

- * Embedded Topic: Metric System Review
- * Embedded Topic: Mass, Volume, and Density
- * Embedded Topic: Significant Figures
- * Embedded Topic: Calculate % Error
- * Embedded: Determine the appropriate tools to gather precise and accurate data.
- * Embedded: Evaluate the accuracy and precision of data.
- * Laboratory Investigation: Density Lab
- * Laboratory Investigation: Counting atoms
- * Embedded: Analyze the components of a properly designed scientific investigation.
- * Embedded: Determine why a conclusion is free of bias.
- * Laboratory Safety

Unit 2: Matter and Energy

Chem1.PS1.5 Conduct investigations to explore and characterize the behavior of gases (pressure, volume, temperature), develop models to represent this behavior, and construct arguments to explain this behavior. Evaluate the relationship (qualitatively and quantitatively) at STP between pressure and volume (Boyle's law), temperature and volume (Charles's law), temperature and pressure (Gay-Lussac law), and moles and volume (Avogadro's law), and evaluate and explain these relationships with respect to kinetic-molecular theory. Be able to understand, establish, and predict the relationships between volume, temperature, and pressure using combined gas law both qualitatively and quantitatively.

Chem1.PS1.6 Use the ideal gas law, $PV = nRT$, to algebraically evaluate the relationship among the number of moles, volume, pressure, and temperature for ideal gases.

Laboratory Investigation: Three Station Gas Laws Lab

* Laboratory Investigation: Gas Laws Lab.

Chem1.PS3.1 Contrast the concepts of temperature and heat in macroscopic and microscopic terms. Understand that thermal energy is a form of energy and temperature is a measure of average kinetic energy of a group of particles.

Chem.PS3.2 Draw and interpret heating and cooling curves and phase diagrams. Analyze the energy changes involved in calorimetry by using the law of conservation of energy quantitatively (use of $q = mc\Delta T$) and qualitatively.

Chem.PS3.3 Distinguish between endothermic and exothermic reactions by constructing potential energy diagrams and explain the differences between the two using chemical terms (e.g. activation energy). Recognize when energy is absorbed or given off depending on the bonds formed and bonds broken.

Chem.PS3.4 Analyze energy changes to explain and defend the law of conservation of energy.

* Laboratory Investigation: Heating and Cooling Curve

* Laboratory Investigation: Calorimetry

Unit 3: Atomic Theory and Structure / Nuclear Chemistry

Chem.PS1. 11 Develop and compare historical models of the atom (from Democritus to quantum model) and construct arguments to show how scientific knowledge evolves over time, based on experimental evidence, critique, and alternative interpretations.

Chem.PS1.9 Draw models (qualitative models such as pictures or diagrams) to demonstrate understanding of radioactive stability and decay. Understand and differentiate between fission and fusion reactions. Use models (graphs or tables) to explain the concept of half-life and its use in determining the age of materials (such as radiometric dating).

Chem.PS1.10 Compare alpha, beta, and gamma radiation in terms of mass, charge, and penetrating power. Identify examples of applications of different radiation types in everyday life (such as its applications in cancer treatment).

- * Half-Lives Laboratory
- * Model Radioactive Decay and Stability
- * Reading Assignment Nuclear Applications

Unit 4: Periodic Table

Chem.PS1.12 Explain the origin and organization of the Periodic Table. Predict chemical and physical properties of main group elements (reactivity, number of subatomic particles, ion charge, ionization energy, atomic radius, and electronegativity) based on location on the periodic table. Construct an argument to describe how the quantum mechanical model of the atom (e.g., patterns of valence and inner electrons) defines periodic properties. Use the periodic table to draw Lewis dot structures and show understanding of orbital notations through drawing and interpreting graphical representations (i.e., arrows representing electrons in an orbital).

* Embedded Topic: Calculating protons, neutrons, and electrons from atomic number and atomic mass for a specific isotope.

Chem.PS4.1 Using a model, explain why elements emit and absorb characteristic frequencies of light and how this information is used.

Unit 5: Bonding and Nomenclature

Chem.PS1.13 Use the periodic table and electronegativity differences of elements to predict the types of bonds that are formed between atoms during chemical reactions and write the names of chemical compounds, including polyatomic ions using the IUPAC criteria.

Chem.PS1.14 Use Lewis dot structures and electronegativity differences to predict the polarities of simple molecules (linear, bent, trigonal planar, trigonal pyramidal, tetrahedral). Construct an argument to explain how electronegativity affects the polarity of basic chemical molecules.

* Laboratory: Use Chemical and Physical Properties to predict if a substance is ionic or molecular.

* Laboratory: Model Molecular Geometry

Chem.PS2.1 Draw, identify, and contrast graphical representations of chemical bonds (ionic, covalent, and metallic) based on chemical formulas. Construct and communicate explanations to show that atoms combine by transferring or sharing electrons.

Chem.PS2.2 Understand that intermolecular forces created by the unequal distribution of charge result in varying degrees of attraction between molecules. Compare and contrast the intermolecular forces (hydrogen bonding, dipole-dipole bonding, and London dispersion forces) within different types of simple substances (only those following the octet rule) and predict and explain their effect on chemical and physical properties of those substances using models or graphical representations.

Unit 6: Chemical equations and Reactions

Chem.PS1.2 Demonstrate that atoms, and therefore mass, are conserved during a chemical reaction by balancing chemical equations.

Chem.PS1.4 Use the reactants in a chemical reaction to predict the products and identify reaction classes (synthesis, decomposition, combustion, single replacement, double replacement).

* Embedded: Writing Chemical Equations, See CHEM1.PS1.13

* Laboratory Investigation: Chemical Reaction Lab.

Unit 7: Stoichiometry

Chem.PS1.1 Understand and be prepared to use values specific to chemical processes: the mole, molar mass, molarity, and percent composition.

Chem.PS1.3 Perform stoichiometric calculations involving the following relationships: mole-mole; mass-mass; mole-mass; mole-particle; and mass-particle. Show a qualitative understanding of the phenomenon of percent yield, limiting, and excess reagents in a chemical reaction through pictorial and conceptual examples. (states of matter liquid and solid; excluding volume of gasses)

* Laboratory Investigation: Stoichiometry

Unit 8: Solutions and Acid/Base

Chem.PS1.7 Analyze solutions to identify solutes and solvents, quantitatively analyze concentrations (molarity, percent composition, and ppm), and perform separation methods such as evaporation, distillation, and/or chromatography and show conceptual understanding of distillation. Construct an argument to justify the use of certain separation methods under different conditions.

Chem.PS1.15 Investigate, describe, and mathematically determine the effect of solute concentration on vapor pressure using the solute's van 't Hoff factor on freezing point depression and boiling point elevation.

Chem.PS2.3 Construct a model to explain the process by which solutes dissolve in solvents, and develop an argument to describe how intermolecular forces affect the solubility of different chemical compounds.

Chem.PS2.4 Conduct an investigation to determine how temperature, surface area, and stirring affect the rate of solubility. Construct an argument to explain the relationships observed in experimental data using collision theory.

Chem.PS1.8 Identify acids and bases as a special class of compounds with a specific set of properties.

****ACT standards are embedded into ALL UNITS throughout the year. Scientific investigations will be conducted for each standard that addresses interpreting data from multiple sources while evaluating models, inferences, and experimental results.****